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- a) Limewater is the common name for a diluted solution of calcium hydroxide
- c) Range 1-3 d) Strong Base is  $\text{NaOH}$  and Weak Base is  $\text{NH}_4\text{OH}$
- e) A hydrated salt is copper sulphate

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crystals ( $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ). f) A universal indicator is a mixed indicator of organic chemicals which not only shows whether the given solution is acidic or basic, but also shows the approximate pH values by giving a wide ...

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Fill in the blanks (a)  
The most commonly  
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is \_\_\_\_\_ and is extracted  
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1. acids
2. citric acid
3. Base
4. calcium sulphate hemihydrate,  $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$
5. yellow colour, pink colour
6. Hydrochloric Acid (HCl)
7. Acetic Acid ( $\text{CH}_3\text{COOH}$ )
8. Sodium Hydroxide ( $\text{NaOH}$ )
- 9.

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Ammonium Hydroxide  
(NH<sub>4</sub> OH) 10. sodium  
carbonate, —Na<sub>2</sub> CO  
3.10H<sub>2</sub> O, permanent  
hardness 11. PH 12.  
Acid 13. Base 14.  
Neutral 15. decreases  
16 ...

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Acid + Base  $\rightarrow$  Salt +  
Water Since, the  
reaction between acid  
and base both

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neutralize each other, hence, it is also known as Neutralization Reaction. Examples: Sodium chloride and water are formed when hydrochloric acid reacts with sodium hydroxide (a strong base).

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Science Chapter 2 -  
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Neutralization is the reaction between acid and base to form salt and water only. From ionic equation, we know that the reaction only involves  $H^+$  ions from acids with  $OH^-$  ions from alkali to form

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water . E.g.  $\text{NaOH} + \text{H}_2\text{SO}_4$  forms  $\text{Na}_2\text{SO}_4$   
+  $\text{H}_2\text{O}$

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substance that produces hydrogen ions ( $H^+$ ) when dissolved in water, whereas a base is a substance that produces hydroxide ions ( $OH^-$ ) in

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10. Explain in terms of the partial charge on hydrogen why NaOH is a base, HClO is a weak acid and HClO<sub>4</sub> is a

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strong acid. 11. Why is HCl a strong acid and HClO a weak acid? 12. Why are HCl and HClO<sub>4</sub> both strong acids? 13. For each of the reactions below, classify the reactants as an acid or a base and the products as the conjugate acid or ...

**Acids and Bases 1  
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