

Chapter 9 Stoichiometry Section 2 Worksheet

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Chapter 9 Stoichiometry Section 2

Stoichiometry is the part of chemistry that applies the balanced chemical equation to determine the quantities of reactants and products. Interpreting balanced equations. ... Chapter 9: Section 2: Ideal Stoichiometric Calculations Last modified by: Michelle Stover ...

Chapter 9: Section 2: Ideal Stoichiometric Calculations

9-2 Ideal Stoichiometric Calculations Ideal Stoichiometry - All reactants are converted into products I. A Common Method for Solving All Stoichiometry Problems A. Mass-Mass Problems 1. Start with a known mass of reactant or product, find an unknown mass of another reactant or product 2. All other stoichiometry problems are derivations ...

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation ...

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CHAPTER 9 REVIEW

Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

93.9% if the percentage yield for the reaction represented by the following equation is calculated to be 75.3%, what mass of Al is expected from the reaction of 52.5g of Al_2O_3 ? $2\text{Al}_2\text{O}_3(\text{l}) \rightarrow 4\text{Al}(\text{s}) + 3\text{O}_2(\text{g})$

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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Chapter 9 Intro to Stoichiometry Section 9.1 . Chapter 9 9.1 Objectives • Define stoichiometry. • Describe the importance of the mole ratio in stoichiometric calculations. • Write a mole ratio relating two substances in a chemical equation. ... Chapter 9 9.2 Objectives

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